

Chapter 8 Covalent Bonding Test A Answers Diantiore

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a bond formed by the sharing of electrons between atoms a covalent bond in which one atom contributes both bonding electrons a molecule consisting of two atoms intermolecular forces resulting from attraction of oppositely charged regions of polar molecules a molecule that has two poles or regions with opposite charges

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a chemical bond in which only one pair of electrons is shared by two bonded atoms structural formula a chemical formula that shows the arrangement of atoms in a molecule or a polyatomic ion unshared pair a pair of valence electrons that is not shared between atoms lone pair double covalent bond bond involving shared pairs

of electrons

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single bonds are always sigma double bonds are 1 pi and 1 sigma and triple bonds are 2 pi and 1 sigma intermolecular forces forces of attraction or repulsion which act between neighboring particles atoms molecules or ions

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a formula that represents the covalent bonds as dashes and shows the arrangement of covalently bonded atoms

unshared pair a pair of valence electrons
that is not shared between atoms double
covalent bond a bond that involves two
shared pairs of electrons triple covalent
bond

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